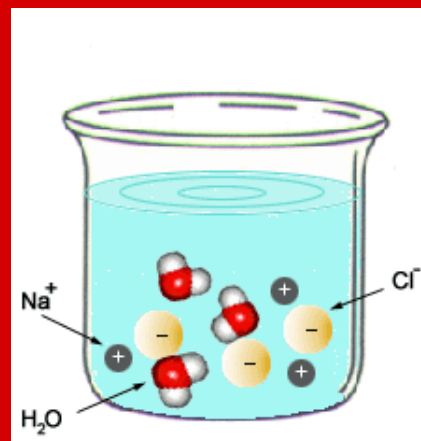


WRITE THIS:

Neutralization



objectives

Review naming acids and bases.

Recognize neutralization reactions

Review law of conservation of mass:

predicting products and balancing reactions

Demos:

cabbage juice ---neutralization reaction

pH probe ---

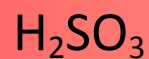
MOM demo

Review

What are the rules for naming binary acids?

What are the rules for naming oxyacids?

Name these acids:



Name these bases:



Write these formulas

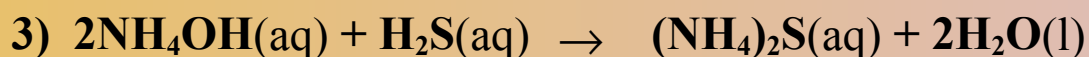
perchloric acid

chloric acid

hydrochloric acid

chlorous acid

Which reactions are double replacement? **Why?**



Which reactions are neutralization of acid and base?

How do you know?

Neutralization reactions

- 1. are a type of double replacement reaction**
- 2. reactants must be ACID and BASE**

What are the products?

answers:

double replacement reactions: **# 1,2,3**

neutralization reactions: **#1 and 3**

Neutralization reactions

1. are a type of double replacement reaction
2. reactants must be ACID and BASE
3. the products are WATER and SALT.

(an ionic compound)



another example:



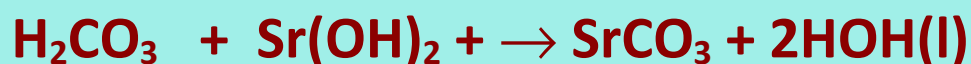
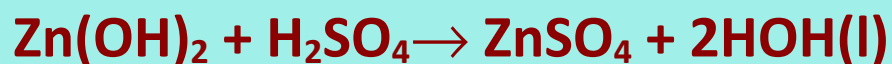
Write the reactant formulas, then predict the products and balance

Zinc hydroxide and sulfuric acid →

potassium hydroxide + chlorous acid →

carbonic acid and strontium hydroxide →

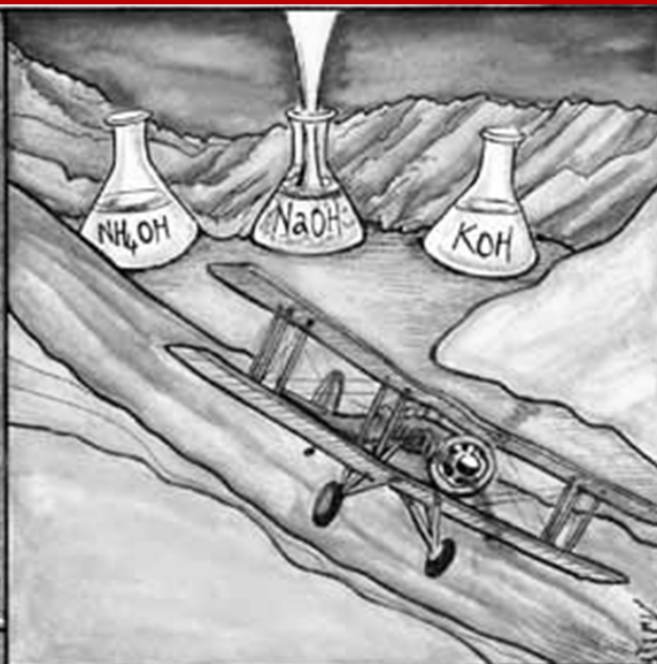
Answers:



Neutralization



DESPITE THE HEAVY FLAK, M'ALISTER'S AIM WAS TRUE, AND HIS CAREFULLY MEASURED ALIQUOT OF HYDROCHLORIC ACID FOUND ITS MARK DEEP IN THE ENEMY'S RESERVOIR OF SODIUM HYDROXIDE.



M'ALISTER GRINNED WRYLY: FINALLY, ONE OF THE ENEMY'S STRONGEST BASES HAD BEEN COMPLETELY NEUTRALISED.

more practice

iron(II) hydroxide + phosphoric acid \rightarrow

ammonium hydroxide + hydrosulfuric acid \rightarrow

answers:

